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Concepts: Terms in this  
set (57) Dalton. The

atom is a tiny,  
indestructible particle  
with no internal  
structure. Thomson.

The atom is a sphere of  
positive electrical  
charge with electrons  
embedded in the  
sphere.

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*Page 5/24*

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CHAPTER 5 What You'll

Learn You will compare

the wave and particle

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models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons in atoms through orbital

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in Atoms. Chapter 5 -  
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Jennie L. Borders.

Section 5.1 - Models of  
the Atom. The

Rutherford's model of  
the atom did not

explain how an atom

can emit light or the

chemical properties of

an atom. Plum Pudding

Model Rutherford's

Model. The Bohr Model.



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## **Electrons in Atoms**

...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how:

Aufbau principle - electrons enter the lowest energy first.

This causes difficulties because of the overlap of orbitals of different energies - follow the diagram! Pauli

Exclusion Principle - at most 2 electrons per

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orbital ...

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Chapter 5 - electrons in  
atoms (handouts)

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table & trends

(handouts) Chapters

7/9 - ionic bonding &  
naming (handouts)

Chapters 8/9 - covalent  
bonding & chemical

names & formulas

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- moles (handouts)

Chapter 11 - reactions

(handouts) Chapter 12

...

**Science / Chapter 5 -  
electrons in atoms  
(handouts)**

Chapter 5: Electrons in  
Atoms Models of the  
Atom Rutherford used  
existing ideas about  
the atom and proposed  
an atomic model in  
which the electrons

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Key  
move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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brinkley504. Ms.

Cristina Chemistry I H.

Terms in this set (82)

Rutherford's model-

also called nuclear

model-did not explain

how the electrons were

arranged

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Section 5.1 - Models of

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Key  
Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

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Section 55.1.1 Figure

5.1 Different elements

can have similar

reactions with water.

Objectives Compare

the wave and particle

natures of light. Define

a quantum of energy,

and explain how it is

related to an energy

change of matter.

Contrast continuous

electromagnetic

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IB Chapter Questions.  
02:25. Problem 1

Objects get their colors  
from reflecting only  
certain wavelengths  
when hit with white  
light. Light reflected  
from a green leaf is  
found to have a  
wavelength of  $4.90 \times 10^{-7}$   
 $\text{m}$ . What  
is the frequency of the  
light?

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Matter an...**

How are electrons  
arranged in an atom?  
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times. 0. 10th grade .

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Chapter 5“Electrons in Atoms ... Describe the energies and positions of electrons according to the quantum mechanical model.

Section 5.1Models of the Atom. OBJECTIVES:

Describe how the shapes of orbitals related to different sublevels differ. Ernest

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How many electrons  
can each p orbital

hold? Chapter 5:

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